

INFLUENCE OF THE Al CONTENT ON THE PHASE TRANSFORMATIONS IN Cu-Al-Ag ALLOYS

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Abstract

The influence of the Al content on the phase transformations in Cu-Al-Ag alloys was studied by classical differential thermal analysis (DTA), optical microscopy (OM) and X-ray diffractometry (XRD). The results indicated that the increase in the Al content and the presence of Ag decrease the rate of the β_1 phase decomposition reaction and contribute for the raise of this transition temperature, thus decreasing the stability range of the perlitic phase resulted from the β_1 decomposition reaction.

Keywords: phase transitions; martensitic transformation; Cu-Al-Ag alloys.

Introduction

The literature about alloys of the Cu-Al system is extensive, due to its practical use. These alloys are of low cost and present good strength and hardness, depending on the Al content. They are used for making sleeves, guides, seats, flanges, gear wheels and many other critical small parts, as electrolysis cell cathodes, circuit plates and casting moulds when high thermal conductivity is needed. Cu-Al alloys with concentration in the range from 9 to 14wt.%Al contain in their structure the ($\alpha + \gamma_1$) eutectoid, which is formed through decomposition of the β phase. On rapid cooling from high temperatures, the β phase undergoes a martensitic transformation. When quenched from the β state, alloys containing more than about 11wt.%Al first become ordered to the β_1 phase and then transform martensitically to either β_1' , ($\beta_1' + \gamma'$) or γ' ,

depending on the Al content. In alloys with less than 11wt.%Al the β phase transforms to an fcc type of disordered martensite [8]. Silver additions to the Cu-Al alloys increase its hardness [4] and improves the resistance to stress corrosion [1], without changing significantly the plasticity and workability of the alloys.

In this work, the influence of the Al content on the phase transformations of Cu-Al-Ag alloys was studied by classical differential thermal analysis (DTA), optical microscopy (OM) and X-ray diffractometry (XRD), and the results were compared with those found in the literature^[2]

Experimental

The Cu-9wt.%Al-4wt.%Ag, Cu-10wt.%Al-4wt.%Ag and Cu-11wt.%Al-4wt.%Ag alloys were prepared in an induction furnace under

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argon atmosphere using 99.95% starting materials and graphite crucible. Results from chemical analysis indicated a final alloy composition very close to the nominal one. Small cylinders of about 10 mm length and 5.0 mm diameter were used for DTA analysis and flat square samples of about 10 mm were obtained for metallography and XRD. These samples were initially annealed for 120 h at 850 °C for homogenization and after annealing some of them were equilibrated for one hour at 850 °C and quenched in iced water. The XRD diagrams were obtained using a Siemens D5000 X-ray diffractometer. After the heat treatments the flat samples were polished, etched and examined in a Leica DMR optical microscope.

DTA curves were obtained using a sample holder with two Ni/Cr-Ni/Al thermocouples, one of which was introduced in the sample and the other in a pure copper cylinder with the same sample dimensions. The sample holder, inside a Vycor tube, was introduced in a furnace and the thermocouples terminals were connected to a HP 34404A multimeter. The heating rate and the data acquisition were controlled using a MQ 112 Micro-Química interface.

Results and discussion

Figure 1 shows the DTA curves obtained at a heating rate of 20 °Cmin⁻¹ for the Cu-9wt.%Al-4wt.%Ag (curve a), Cu-10wt.%Al-4wt.%Ag (curve b) and Cu-11wt.%Al-4wt.%Ag (curve c) alloys annealed at 850 °C for 120 hours. In curves (a) and (b) five endothermic peaks were observed, while in curve (c) only three peaks were detected. Peak P₁, at about 559 °C, is due to the $\alpha + \gamma_1 \leftrightarrow \beta$ eutectoid transformation [10], as expected from the Cu-Al equilibrium diagram [9]. The endothermic peak P₂, at about 508 °C, is associated to the β_1 phase disordering. The martensitic phase β_1' changes into the β_1 phase in the same temperature interval of the α_2 phase disordering process and at about 508 °C the β_1 phase transforms into the β phase [7]. Peak P₃, at about 380 °C, is associated to the dissolution of the α_2 phase precipitates formed during slow cooling [4]. The peak P₄, which is observed at about 330 °C in curve (a) and 270 °C in curve (b), corresponds to the $\alpha + \alpha_2 \rightarrow \alpha + \gamma_1$ order-disorder transition⁶ and peak P₅, observed at about 220 °C in curve (a) and 170 °C in curve (b)

corresponds to the beginning of the α_2 phase disordering [2]. Peaks P₄ and P₅ were not detected in curve (c). It is possible to observe, in fig.1, that the intensity of peak P₁ increases with the Al concentration and that its temperature is shifted to lower values, when compared with the temperature for samples without Ag addition [2]. This temperature shift may be associated to the Ag dissolution in the matrix, which increases the α phase relative fraction and facilitates the formation of the β phase from the combination of the remaining α with the perlitic phase. The increase in the intensity of peak P₁ may be due to the increase in the Al concentration, which causes an increase in the relative fraction of the γ_1 phase.

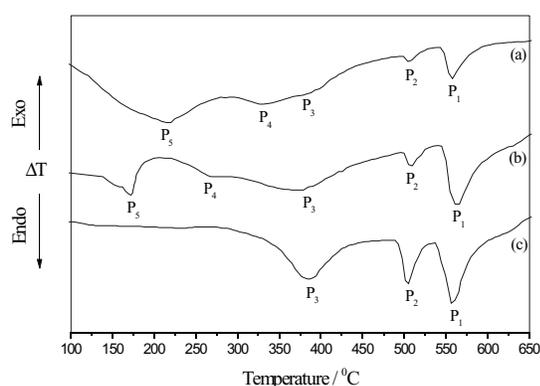


Figure 1. DTA curves obtained after annealing: (a) Cu-9wt.%Al-4wt.%Ag alloy; (b) Cu-10wt.%Al-4wt.%Ag alloy and (c) Cu-11wt.%Al-4wt.%Ag alloy. Heating rate 20 °C min⁻¹.

Peak P₂ is also shifted to lower temperatures, compared with samples without Ag additions, and its intensity increases with Al concentration. The shift and intensity increase may be related to the combined effects of Al concentration increase, which facilitates the β_1 phase rosettes formation [5] and the aftereffect of β phase formation from β_1 , and of the presence of Ag, which seems to increase the relative fraction of the β_1' martensitic phase retained in samples cooled at a rate higher than 2 °Cmin⁻¹. Peak P₃ intensity increases with the increase in the Al content, due to the greater α_2 relative fraction. The presence of Ag seems not to interfere in this transformation, because it is observed at the same temperature as the samples without Ag. Peaks P₄ and P₅ are shifted to lower temperatures with the increase in the Al concentration, and for the Cu-11wt.%Al-4wt.%Ag (curve c) they are not detected.

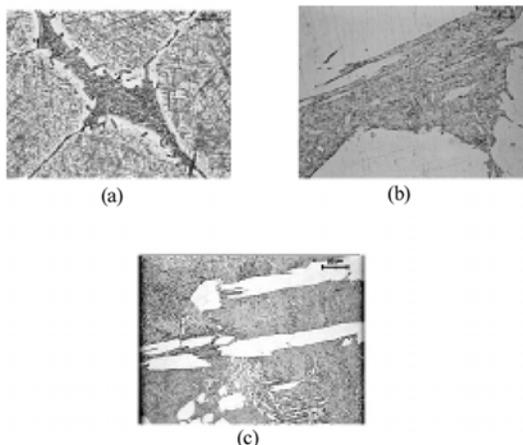


Figure 2. Optical micrographs (500x) obtained for the alloys after annealing: (a) Cu-9%Al-4%Ag; (b) Cu-10%Al-4%Ag; (c) Cu-11%Al-4%Ag.

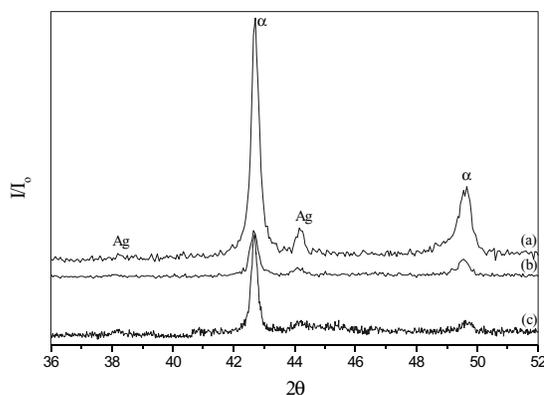


Figure 3. X-ray diffraction patterns obtained for the alloys after annealing: (a) Cu-9%Al-4%Ag; (b) Cu-10%Al-4%Ag; (c) Cu-11%Al-4%Ag.

The starting point of the DTA curves in fig. 1 corresponds to the perlitic phase, as observed in the optical micrographs of figure 2. At this point there is also an amount of Ag remaining from the solution treatment, as showed in the X-ray diffraction patterns of figure 3. The micrographs of fig. 2 seems to confirm that the increase in the Al content decreases the relative fraction of α phase and promotes the perlitic phase formation.

Figure 4 shows the DTA curves obtained at a heating rate of 20°Cmin^{-1} for the Cu-9wt.%Al-4wt.%Ag (curve a), Cu-10wt.%Al-4wt.%Ag (curve b) and Cu-11wt.%Al-4wt.%Ag (curve c) alloys

quenched from 850°C . In the curves of fig. 4 the endothermic peak P_1 , at about 559°C , is due to the $\alpha + \gamma_1 \leftrightarrow \beta$ eutectoid transformation [10] and the endothermic peak P_2 , at about 525°C , is attributed to the $\beta_1 \rightarrow \beta$ transition from the remaining part of the β_1 phase formed at low temperature [8]. The endothermic peak P_3 , at about 480°C , is associated to the reverse martensitic transformation and to the $\beta_1 \rightarrow \alpha + \gamma_1$ decomposition from part of the β_1 phase. Peak P_3 was not detected in curves (a) and (b). The exothermic peak P_4 , at about 400°C , may be associated to the precipitation of Ag dissolved in the matrix [3]. P_5 is due to the ordering of the martensitic β' phase remaining from the $\beta' \rightarrow \beta_1'$ transformation and to the end of the α_2 phase ordering [2]. P_6 may be associated to the beginning of the α phase ordering. Peak P_1 in fig. 4 seems to undergo the same changes observed for the corresponding peak in fig. 1. It is possible to observe an increase in the intensity of peak P_2 in fig. 4, but the transition temperature is very close to the expected one [2]. Comparing to the transition for the annealed sample (fig. 1), this phase transition is now the dominant effect for the quenched samples, in the temperature range considered. This must be associated to the increase in the martensitic phase retained at fast cooling. This peak temperature is close to that observed for samples without Ag addition, indicating that the effect of the presence of Ag has been minimized by the heat treatment. Peak P_3 , observed at about 480°C in curve (c) of fig. 4, is observed at about 380°C in Cu-Al alloys [2], indicating that the increase in the Al concentration and the presence of Ag are decreasing the β_1 phase decomposition reaction rate and contributing to the increase in the temperature transition. Thus it causes a decrease on the stability range of the perlitic phase produced by the β_1 phase decomposition. As the $\beta_1' \rightarrow \beta_1 \rightarrow \beta$ transition and the β_1 phase decomposition occur in the same temperature interval, it is also observed an increase in the stability range of the martensitic phase with the Al concentration and the presence of Ag. Peak P_4 seems to indicate that the beginning of the β_1 phase decomposition may be related to Ag precipitation. It is known that Ag diffusion in the γ_1 phase is greater than in the α phase, due to changes in the atomic density [11]. In this way, Ag precipitation must occur mainly on the perlitic phase and this seems to indicate that

the β_1 phase decomposition catalyzes the Ag precipitation reaction. The absence of the exothermic peak P_4 in curve (a) seems to indicate that this thermal event is occurring in the same temperature range of the endothermic peak P_3 , and that the enthalpic variations are similar, causing the absence of both thermal events. Peak P_5 is observed at about 290 °C in alloys with 9 and 10 wt.%Al and is associated to the martensitic ordering reaction [2] and is not observed in curve (a) of fig. 4, indicating that the presence of Ag disturbs the β' phase ordering process. In curve 4-b this peak was observed at about the same temperature as that for the alloy without Ag addition.

The starting point of the DTA curves of fig. 4 corresponds to the martensitic phase, as showed by the micrographs in figure 5. It is known [12] that alloys with less than 10.8 wt.%Al must show the β' type martensitic phase and that over this concentration the martensitic phase observed is the β_1' . The micrographs in fig. 5 show the presence of the β_1' martensitic phase in the Cu-10 wt.%Al alloy. It seems to indicate that the addition of 4 wt.%Ag shifts the equilibrium concentration to higher Al contents. The presence of the β_1' martensitic phase in the Cu-10 wt.%Al-4 wt.%Ag alloy indicates that the dominant effect, in the temperature range where peak P_5 is detected, is the end of the α phase ordering process. In curve 4-c this final step is shifted to higher temperatures, probably due to some interaction of the α phase with Ag.

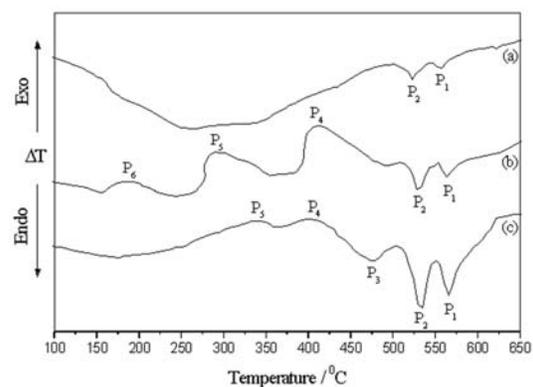


Figure 4. DTA curves obtained for alloys initially quenched: (a) Cu-9%Al-4%Ag; (b) Cu-10%Al-4%Ag; (c) Cu-11%Al-4%Ag. Heating rate 20 °Cmin⁻¹

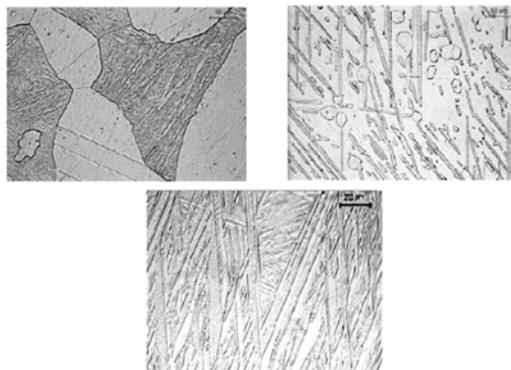


Figure 5. Optical micrographs (500x) obtained for alloys quenched from 850 °C: (a) Cu-9%Al-4%Ag; (b) Cu-10%Al-4%Ag; (c) Cu-11%Al-4%Ag.

In the micrographs of figure 6 it is possible to observe the phase transformations sequence occurring in the Cu-10 wt.%Al-4 wt.%Ag alloy initially quenched from 850 °C and then successively quenched in the temperature range from 300 to 950 °C. These micrographs seem to confirm what was proposed in the discussions of fig. 4. In the X-ray diffraction patterns showed in figure 7 it is possible to observe that the proposed sequence is in good agreement to that expected from the Cu-Al equilibrium diagram [9], and that this sequence of transformations is not modified by the presence of Ag.

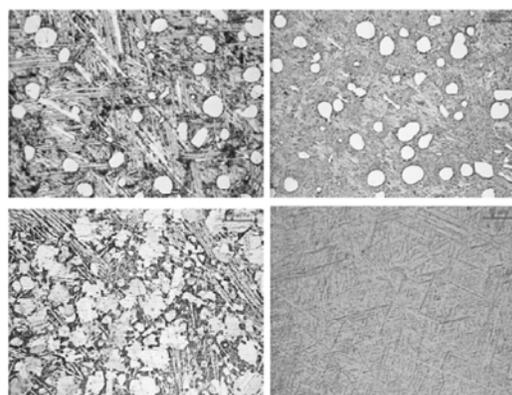


Figure 6. Optical micrographs (200x) obtained for the Cu-10%Al-4%Ag alloy initially quenched from 850 °C and then quenched from: (a) 300 °C; (b) 400 °C; (c) 600 °C and (d) 950 °C.

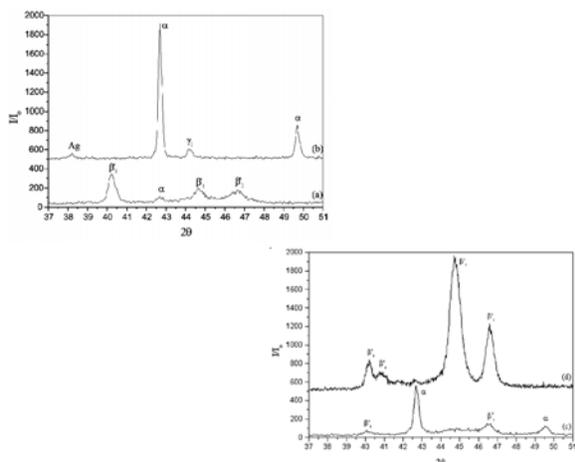


Figure 7. X-ray diffraction patterns obtained for the Cu-10%Al-4%Ag alloy initially quenched from 850 °C and then quenched from: (a) 300 °C; (b) 500 °C; (c) 600 °C and (d) 950 °C.

Conclusions

The results indicate that with the addition of 4 wt.%Ag to the Cu-9wt.%Al, Cu-10wt.%Al and Cu-11wt.%Al alloys it is possible to observe the phase transformations corresponding to higher Al concentration. The increase in the Al content decreases the α phase relative fraction and promotes the perlitic phase formation. The increase in the perlitic phase seems to facilitate the Ag diffusion. The increase in the Al concentration and the presence of Ag decrease the β_1 phase decomposition reaction rate and contribute for the increase in the temperature transition, thus decreasing the stability range of the perlitic phase, produced by the β_1 phase decomposition reaction. The results also indicate that this decomposition reaction seems to catalyze Ag precipitation.

Acknowledgements

The authors thank FAPESP and CNPq for financial support.

ADORNO, A. T.; BENEDETTI, A. V.; SILVA, R. A. G.; BLANCO, M.; Influência da concentração de Al nas transformações de fase em ligas de Cu-Al-Ag.

Resumo

A influência da concentração de Al nas transformações de fase em ligas de Cu-Al-Ag foi estudada por análise térmica clássica (ATD), microscopia óptica (MO) e difratometria de raios X (DRX). Os resultados indicaram que o aumento no teor de Al e a presença de Ag diminuem a velocidade da reação de decomposição da fase β_1 e contribuem para o aumento da temperatura desta transição, diminuindo assim o intervalo de estabilidade da fase perlítica, resultante da reação de decomposição.

Palavra-chave: transformações de fase; transformação martensítica; ligas de Cu-Al-Ag

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